by said subsequent neutralization to a value of at least 4.5~~, preferably at least 5~~.

4. (Currently Amended) Method according ~~to any one of Claims 1 to 3, characterized in that~~ Claim 3, wherein said preliminary and subsequent neutralizations are carried out using a strong base selected from the group consisting of the hydroxide, the oxide and the water-soluble salts of calcium, sodium, potassium and/or ammonium.

5. (Currently Amended) Method according to ~~any one of Claims 1 to 4, characterized in that~~ Claim 4, wherein the digestion is carried out at ambient temperature.

6. (Currently Amended) Method according to ~~any one of Claims 1 to 5, characterized in that~~ Claim 5, further comprising ~~it comprises~~ a preliminary step of forming said aqueous solution of hydrochloric acid by diluting concentrated hydrochloric acid in water.

7. (Currently Amended) Method according to ~~any one of Claims 1 to 6, characterized in that~~ Claim 6, further comprising ~~it comprises~~ a preliminary step of forming said aqueous solution of hydrochloric acid by treating an aqueous solution of calcium chloride with sulphuric acid and removing a calcium sulphate precipitate therefrom.

8. (Currently Amended) Method according to ~~any one of Claims 1 to 7, characterized in that~~ Claim 7, wherein the phosphate ore has a ~~P2O5~~ P_2O_5 content of 25 to 35% by weight.

9. (Currently Amended) Method according to ~~any one of Claims 1 to 8, characterized in that~~ Claim 8, wherein said aqueous solution of hydrochloric acid which is used in the digestion has an HCl concentration of around 3 to 8%~~, preferably 5 to 7.4%~~ by weight.

10. (Currently Amended) Method according to ~~any one of Claims 1 to 9, characterized in that~~ Claim 9, further comprising ~~it comprises~~ a treatment of said aqueous solution of calcium chloride with an aqueous solution of sulphuric acid, with formation of insoluble calcium sulphate, which precipitates, and of an aqueous phase based on hydrochloric acid, an isolation of the calcium sulphate precipitate, and an at least partial recycling, to the digestion step, of the aqueous phase based on hydrochloric acid, so as to form said aqueous solution of hydrochloric acid.

11. (Currently Amended) Method according to ~~any one of Claims 1 to 9, characterized in that~~ Claim 9, further comprising ~~it also comprises~~ an additional neutralization of said aqueous solution of calcium chloride, so as to adjust this aqueous solution to a pH which is greater than the pH of the subsequent neutralization and so as to precipitate residual impurities, and an elimination of these impurities from said aqueous solution, a treatment of the latter with an aqueous solution of sulphuric acid, with formation of insoluble calcium sulphate, which precipitates, and of an aqueous phase based on hydrochloric acid, an isolation of the calcium sulphate precipitate, and a recycling, to the digestion step, of the

aqueous phase based on hydrochloric acid, so as to form said aqueous solution of hydrochloric acid.

12. (Currently Amended) Method according to Claim 11, ~~characterized in that~~ wherein the pH of said aqueous solution of calcium chloride is adjusted by said additional neutralization to a value of at least 8.5~~, preferably at least 9~~.

13. (Currently Amended) Method according to ~~either one of Claims 11 and 12, characterized in that~~ Claim 12, wherein the additional neutralization is carried out using a strong base selected from the group consisting of the hydroxide, the oxide and the water-soluble salts of calcium, sodium, potassium and/or ammonium.

14. (Currently Amended) Method according to ~~any one of Claims 1 to 13, characterized in that~~ Claim 13, wherein the digestion takes place in a reactor equipped with a stirrer.